**2.1 - Practice Time!**

Draw plausible Lewis structures for the following molecules. Tips for drawing structures can be found on the opposite side of this page.

1. Ca
2. S-2
3. H-
4. Sn+2
5. C2N2
6. ICl
7. OF2
8. Br2
9. CS2
10. Cl2CO
11. Ca(OH)2
12. NH4Br
13. Ca(OCl)2

**Tips for Drawing Lewis Structures**

1. Determine the total number of valence electrons that must appear in the structure.
2. Identify the central atom(s) and terminal atoms.
3. Write a plausible skeletal structure. Join the atoms in the skeletal structure by *single* covalent bonds (single dashes).
4. For each bond in the skeletal structure, subtract *two* from the total number of valence electrons.
5. With the valence electrons remaining, *first* complete the octets of the terminal atoms. *Then*, to the extent possible, complete the octets of the central atom(s). If there are just enough valence electrons to complete octets for all of the atoms, the structure is a satisfactory Lewis structure.
6. If one or more central atom is left with an incomplete octet after Step 5, move lone-pair electrons from terminal atoms to form *multiple* covalent bonds to central atoms. Do this to the extent necessary to give all atoms complete octets, thereby producing a satisfactory Lewis structure.